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POWER LAW BASICS | 6.2 POWER LAW BASICS 2.1 ARRHENIUS EQUATION The main assumption behind the Arrhenius expression is that $k = A \cdot e^{-E_a/RT}$. This is an approximation, but it works quite well. The rate coefficient is the term that is a function of temperature but may also depend

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Chemical Equilibrium $dG = -SdT + VdP + \sum_{j=1}^n \mu_j dn_j$ μ_j : chemical potential for species j . $\partial G / \partial \epsilon_i = \sum_{j=1}^n \nu_{ij} \mu_j = 0$, $i = 1, \dots, r$, $\mu_j = G_j + RT \ln a_j$, $a_j = f_j / f_j^0$ $K_i = \prod_{j=1}^n a_j^{\nu_{ij}}$ $\Delta G_i = -RT \ln K_i$ Standard state: pure species j at 1 atm and system temperature.

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